











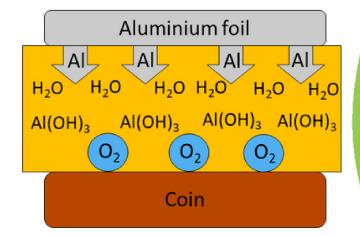


The Aluminium-Air Battery



The Aluminium foil will oxidise into Al³⁺ ions. The electrolyte (salt water) helps to move the charge. The coin acts as an inert conductor for the oxygen reduction reaction that each cell needs. The cloth will hold the electrolyte.

- Aluminium Foil Dish Cloth with Salt Water Coin **Aluminium Foil** Dish Cloth with Salt Water Coin Aluminium Foil Dish Cloth with Salt Water
- 1. Measure and cut out 8 disks of cloth and 8 disks of aluminium.
- 2. Dampen the cloth in the saltwater (not soak)
- 3. Stack the coin, cloth and aluminium



 $4AI + 3O_2 + 6H_2O \rightarrow 4AI(OH)_3$



Coin

www.currentchemistry in vestigators.com

https://ie.rs-online.com/web/

https://rsc.li/takecharge

Hold the *long leg* of the LED to the coin electrode

Hold the *short leg* of the LED to the Aluminium electrode.

What is the fewest number of cells needed?









Cell No.	Salt Water (volts)	Water (volts)	Vinegar (volts)
1			
2			
3			
4			
5			
6			
7			
8			

Potential Difference (Voltage)